

DETERMINING MOLECULAR POLARITY

1. Molecules are **non-polar** or **polar**

2. **Non-polar** molecules have an even (symmetrical) distribution of charge (+ or -)

➤ If all atoms are the same (non-polar bonds)



➤ If there are no lone pairs on the central atom and the attached atoms are all the same



3. **Polar** molecules have an uneven (asymmetrical) distribution of charge. The molecule has a dipole (+ side and a - side, *like a bar magnet*)

➤ If the atoms are different from each other



➤ If there are lone pairs on the central atom and the attached atoms are different from the central atom



Steps for determining molecular polarity in simple covalent compounds:

