

DESCRIPTION OF MATTER (mass and occupies space) **Chp 1.1 & 1.2**

PURE (always homogeneous)	MIXTURE (impure)			
Elements – all atoms are the same Represented by symbols of one element Compounds – made from two or more elements Represented by symbols of two or more elements	homogeneous (uniform throughout – composition/properties)		heterogeneous (not uniform composition/properties)	
	brass	sugar dissolved in water	granite	fog (air water)
	air	tap water	dirt	fresh squeezed lemonade (pulp in it)

STATES OF MATTER - gas, liquid, solid **Chp 1.4**

TRANSFORMATION of matter **Chp 1.6**

<u>Physical – phase change</u> ice melting grinding coffee beans	<u>Chemical- transformation to new substance(s)</u> Hydrogen reacting with oxygen to produce water Hydrogen peroxide decomposing to water and oxygen
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PROPERTIES of matter **Chp 1.6**

<u>Physical (detected using the senses):</u> Boiling point Melting point Shape Color	<u>Chemical (stability):</u> Reactivity with oxygen Stability in air Heat stability
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HOW ARE PROPERTIES AND TRANSFORMATION OF MATTER STUDIED?

- HYPOTHESIS – TENTATIVE EXPLANATION OF AN OBSERVATION THAT CAN BE VERIFIED BY EXPERIMENT. *Experiments must be conducted under carefully controlled conditions so that they can be reproduced over and over. This separates science from pseudo science.*
- THEORY – AN EXPLANATION OF AN OBSERVATION THAT IS SUPPORTED BY EXPERIMENTAL DATA AND WHICH IS USED TO PREDICT OTHER OBSERVATIONS. PHYSICAL OR MATHEMATICAL MODELS ARE FREQUENTLY DERIVED FROM A THEORY. *At any time, a new experiment may be done that invalidates the theory.*
- LAW – A STATEMENT OF OBSERVATION OF A NATURAL PHENOMENA (Law of Conservation of Mass - “The total mass of reactants is equal to the total mass of products in a chemical reaction.”) *Laws do not ever change because they are based on observations of natural phenomena.*